**Activity 3**

**Aim:** To demonstrate the process of electroplating

**Material required/used:** \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_,

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_,\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_,\_\_\_\_\_\_\_\_\_\_\_\_\_\_, \_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Diagram:**

**Procedure:**

1. Take 2 conducting plates – one of Zinc and other of Copper.
2. Take 250mL of distilled water in a beaker and dissolve 2 tablespoons of copper sulphate (CuSO4) in it
3. Add a few drops of dil. Sulphuric Acid (H2SO4) to it to make it more conducting.
4. Connect the copper electrode to +ve terminal and Zinc to the –ve terminal of battery.
5. Allow a steady current to pass through the electrolyte for 30 minutes.
6. Take out the plates from the solution.

**Observation:**

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**Inference:**

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